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calculus are involved  
in learning physical  
chemistry, but they

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are only tools. P-  
chem is so much  
more than just math.  
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there is a brief  
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$\ln K = (\Delta_r S^\circ / R) -$

$(\Delta_r H^\circ / RT)$  The

equilibrium for the

dissociation of one

mole of iodine vapour

is  $3.94 \times 10^{-4}$  at 900K

and the standard



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enthalpy is 154  
kJmol<sup>-1</sup>. Calculate the  
equilibrium constant  
at 1200K assuming  
the standard enthalpy  
remains constant. In  
 $(K_2/K_1) = - \Delta_r H^\circ / R$   
 $(1/T_2 - 1/T_1)$  , ANS =  
 $6.7 \times 10^{-2}$  at 1200K.

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Physical chemistry is the study of how matter behaves on a molecular and atomic level and how chemical reactions occur. Based on their analyses, physical chemists may develop new theories, such as

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how complex structures are formed. Physical chemists often work closely with materials scientists to research and develop potential uses for new materials.

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each. To complete the

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does not have  
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questions. If you want  
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