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$I H P_{vap} = \exp 14.790 ? T K$ and by the ideal gas law f Solutions to Chemical and Engineering Thermodynamics, $3e N V RT \times 18 \text{ mol} \times 8.314 \times 10^7 5$ (bar ? m 3 mol K) $T P = VV 1 \times 10^{73} ? (10 ? x) 106$ volume taken up by liquid Equating P and P vap we have $x 18 \times 8.314 \times 10^{25} T F 5432.8 \times I .$

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